

Synthesis and Applications of Nanocontainers and Nanorattles

Sarah-Luise Abram, Philippe Yep, and Katharina M. Fromm*

Abstract: Nanorattles and hollow nanocontainers have unique properties related to the presence of a void space inside the shell, which has been shown to improve the properties of materials for e.g. catalysis, drug release, sensor materials or lithium ion batteries. This article summarizes our recent progress in the synthesis of hollow silica, titania and ceria nanorattles and nanocontainers, eventually encapsulating AgNPs as cargo and highlights their potential applications in catalysis and biomedicine.

Keywords: Antimicrobial · Catalysis · SiO₂ nanorattles · Silver nanoparticles · TiO₂ nanocontainers



Katharina M. Fromm studied chemistry in Karlsruhe and Strasbourg before receiving a PhD from the TU Karlsruhe in 1994. She carried out her postdoctoral studies with, among others, Nobel-Prize winner Jean-Marie Lehn in Strasbourg, and did her habilitation at the University of Geneva. In 2003, she became SNSF-Professor at the University of Basel before moving to a full professorship at the University of Fribourg in 2006. She is currently Vice-President of the Research Council of the SNSF, Fellow of the American Chemical Society, elected member of the European Academy of Sciences and recipient of the Prix Jaubert of the University of Geneva in 2018.

Hollow Nano/Micro-structures in Nature

A hollow nano/micro-structure is defined as a solid structure with a void space inside a distinct shell and nano- or micrometer dimensions.^[1] Such structures can be classified according to their overall shape, the number of shells, or the presence of one or several movable cores (see Fig. 1A and B); Core and shell materials can be organic, inorganic, or composites, e.g. polymers, ceramics, metals, metal oxides or other chalcogenides.^[2] Nature offers a great variety of such hollow structures exploiting the nano- or microscale dimensions for different purposes. For example, ferritins are large proteins whose structure comprises a hollow cavity used for storage and on demand release of iron (see Fig. 1C).^[3] Pollen microcapsules are used by plants to shield their male sperm from harsh environmental conditions (see Fig. 1D).^[4] Compartmentalization and hierarchical structuring at the micro- or nanoscale are ubiquitous principles in biology realized in membrane-bound structures such as cells, viral envelopes and capsids or other vesicles.^[5] Hence, membrane separation allows the cooperative organization of two otherwise chemically incompatible functionalities as it is e.g. realized for photosystem I and II with the thylakoid membrane which was recently mimicked by polymeric nanocontainers.^[6]

Natural photonic materials are also based on hierarchical hollow nanostructures: The Edelweiss plant uses hollow nanostructured filaments as UV absorbers, that dissipate the energy and thus protect the plant from the damaging radiation.^[7] Another example is the strong iridescence of the hair-like setae of many polychaete

worms (Fig. 1E) resulting from the periodic arrangement of voids into the lattice of a pseudo-photonic crystal.^[8]

Benefits and Applications of Nanocontainers and Nanorattles

Compared to their solid counterparts, hollow nanostructures provide additional properties, which are advantageous for many applications: i) high specific surface area, ii) low density, and iii) high loading capacity for molecular species, gases or nanoparticles.^[11] The low density is interesting for their application in fillers, pigments or coatings^[9] whereas the large surface area is useful for sensing applications^[10] and in catalysis either as the active catalyst^[11] or as support material.^[12] In nanorattles (NR), the movable core particle is protected from aggregation, coalescence, sintering and the outside chemical environment by the shell, thus preserving the core's catalytic activity. Porosity in the shell allows the diffusion of the reactants to the core while the void provides a confined space for the reaction. The NR architecture thus overcomes the problem of a restricted access to the core's active surface in traditional core shell particles.^[12] Moreover, the presence of the void may alter the optical,^[13] magnetic,^[14] electronic or photocatalytic properties of certain materials.^[15] Hollow nanostructures were as well successfully applied in energy storage devices such as Li ion batteries or supercapacitors as their void buffers the huge volume variation during lithiation and delithiation processes^[16] and their high surface area results in a large electrolyte-electrode interface with an increased Li⁺ ion flux.^[17]

More complicated hierarchical hollow structures show a great potential for the development of multifunctional materials, e.g. in theranostics^[18] where imaging agents for diagnosis (fluorescent dyes or particles, but also MRI,^[19] ultrasound^[20] or photoacoustic imaging^[21] contrast agents) are combined with drug delivery or other therapeutic approaches in one product. In this sense, hollow mesoporous silica,^[22] polymeric vesicles,^[23] and different NR^[12] are loaded with anticancer drugs,^[24] ibuprofen,^[25] enzymes,^[26] or siRNA^[27] and are sometimes even modified to achieve a stimuli-responsive drug release activated by temperature,^[26] pH,^[28] enzymes,^[29] or the K⁺ concentration.^[30]

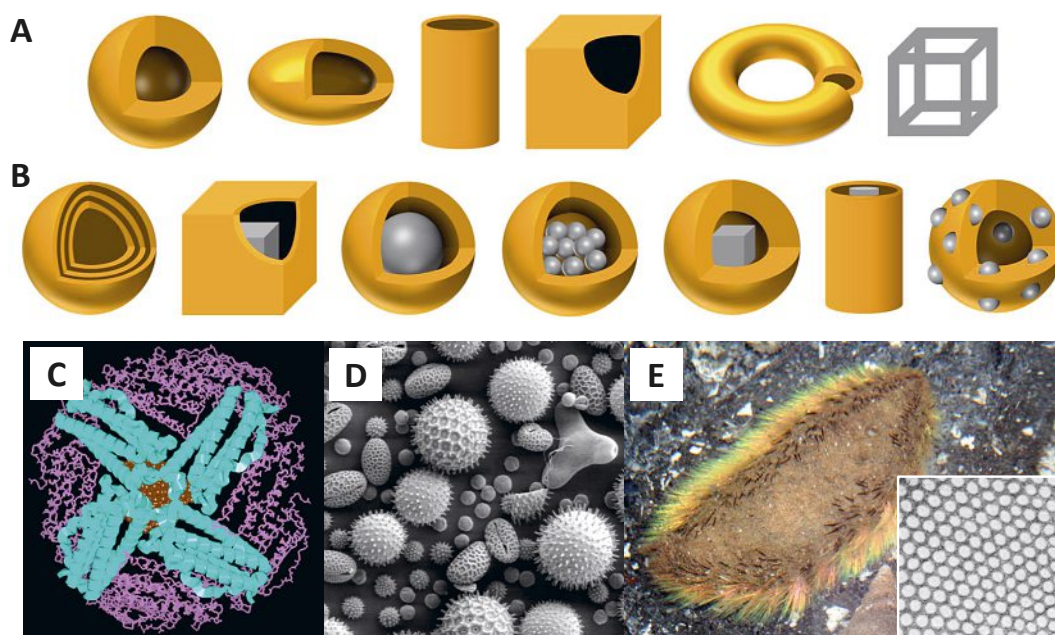
Synthesis of Nanocontainers and Nanorattles

The synthesis of hollow nanostructures can be roughly divided into three approaches for the void creation: (1) hard templating, (2) soft templating, and (3) self-templating.^[1]

In the Fromm group, all three approaches were used for the synthesis of silica, titania and ceria nanocontainers (NC) and nanorattles (NR) encapsulating Ag nanoparticles (AgNPs). The

*Correspondence: Prof. Dr. K. M. Fromm
E-mail: katharina.fromm@unifr.ch
University of Fribourg, Department of Chemistry, Chemin du Musée 9,
CH-1700 Fribourg

Fig. 1. Schematic representation of hollow nanostructures with different shapes and hierarchical levels: (A) hollow spheres, ellipsoids, tubes, boxes/cubes, tori or cages. (B) multi-shelled hollow spheres and cubes (cube-in-box), nano-rattles/yolk-shell particles with one or multiple cores, wire-in tube structures, and decorated hollow nanospheres. (C) Encapsulation of iron inside the ferritin cage, image reprinted from http://www.chm.bris.ac.uk/motm/ferritin/ferritintutorial_molmonth.html (15.11.2018). (D) Electron micrograph of different pollen and (E) Natural photonic crystals in the spines of *Aphrodita* and *Pherusa* (TEM reprinted (adapted) with permission from ref. [8], copyright 2009 American Physical Society, photograph by 'Little Boffin (PeterEdin)' licensed under CC BY-NC-SA 2.0).



scope of these NC and NR ranges from the improvement of Ag-based antimicrobial coatings for medical devices to applications in catalysis for redox reactions, photoelectrochemical water splitting, CO₂ reduction or the degradation of dyes. The following paragraphs will provide an overview of some of the synthesized materials and their applications.

Ag-CeO₂ and Ag-TiO₂ NC by the Hard Template Method

For any hard template syntheses, the selective heterogeneous precipitation of the shell material onto the template as well as a selective removal strategy for obtaining the void without damaging the shell are both crucial for a homogeneous product. Size and shape of the void depend strongly on the template's size and shape, with the shell thickness being controlled by the coating procedure. Common coating strategies comprise sol-gel methods, polymerizations or hydrothermal reactions on polymer, silica, carbon, ceramic, or metal nanoparticles as typical templates.

Since the first report in 1998 by Caruso *et al.* on the synthesis of hollow microstructures using layer-by-layer deposition on polystyrene (PS) templates they became a popular template material for a great variety of shell materials as they are readily available in various diameters and easily removed by dissolution in organic solvents or calcination.^[1,31] Our group has used anionic 200 nm PS spheres for the synthesis of a series of TiO₂ and CeO₂ NC, either with or without the integration of AgNPs or Ag⁺ into the shell and cavity for antimicrobial activity or the improvement of photocatalytic properties (Fig. 2).^[32]

Encapsulation of silver salts inside CeO₂ NC resulted in a sustained release of Ag⁺ as antimicrobial active species over a period of several weeks.^[32b] Replacing the salt by AgNPs enabled a further prolongation of the Ag⁺ release to several months, making Ag-CeO₂ NC promising candidates for long-term infection prevention. The addition of a second shell made from TiO₂ restricted the Ag⁺ release even more while still preserving the antimicrobial activity in agar diffusion tests.^[32a] After three months, the remaining Ag (> 90% of the total initial Ag content) could be released by addition of nitric acid showing the potential for a triggered on-demand release of Ag⁺ from these Ag-CeO₂@TiO₂ NC, as many bacteria tend to acidify their environment.^[33]

The same synthetic route was used for Ag-TiO₂ NC that were tested for the photocatalytic reduction of CO₂ to value-added organic molecules.^[34] A complex hierarchical structure has been shown to enhance the photocatalytic performance of TiO₂ com-

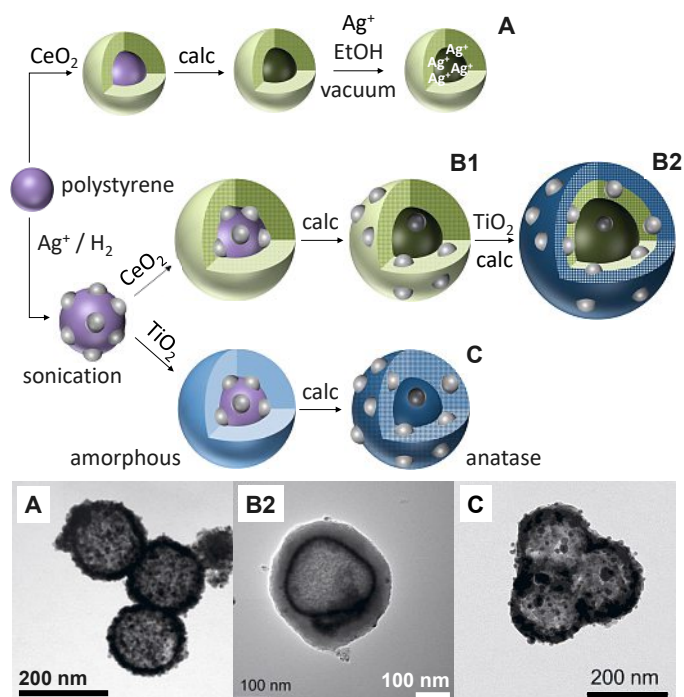


Fig. 2. Scheme of the hard template synthesis of Ag⁺-CeO₂ (A), Ag-CeO₂ (B1), Ag-CeO₂@TiO₂ (B2), and Ag-TiO₂ (C) nanocontainers by sol-gel coating of 200 nm PS spheres (violet) with CeO₂ (green) and TiO₂ (blue) layers and subsequent template removal by calcination and their TEM images. Micrographs (A), (B2) and (C) taken from ref. [32b], ref. [32a] and ref. [34] respectively.

pared to common nanoparticles or bulk materials.^[35] Moreover, integrated AgNPs could act as electron scavengers and thus lower the electron-hole recombination rate which was shown to shift the product mixture to more electron-demanding products at the expense of CO (*i.e.* methane requiring 8 e⁻ instead of 2 e⁻ for CO).^[36] Besides the photoreduction of CO₂, TiO₂ nanomaterials are used for photoelectrochemical water splitting as second contribution to the quest for alternative energy sources and storage.^[37] The NC morphology as well as doping with AgNPs could be interesting ways to increase the photocatalytic efficiency by shifting the absorption to the visible range.^[38]

As a drawback, the above-described Ag-CeO₂ and Ag-TiO₂ NC lacked monodispersity as the integration of AgNPs into the

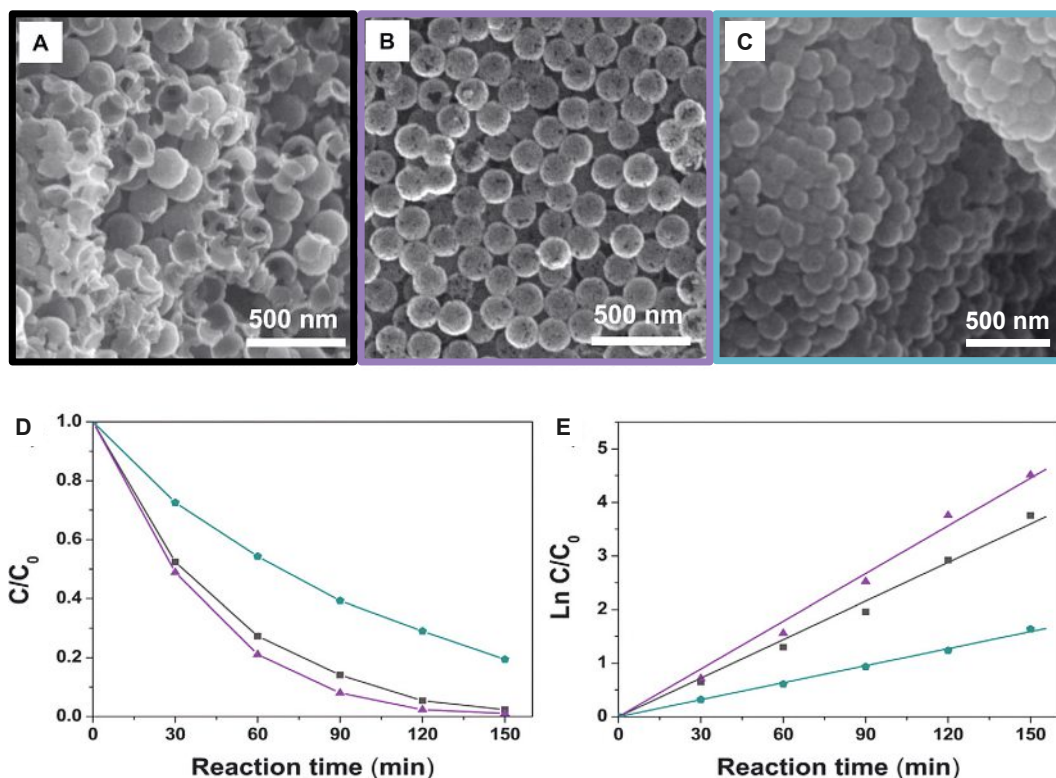


Fig. 3. Scanning electron micrographs of anatase TiO₂ NC by hard template synthesis with (A) direct calcination or (B) Soxhlet extraction prior to calcination and (C) with increased wall thickness. (D) Photodegradation of MB by the mentioned anatase TiO₂ NC and (E) their pseudo first order kinetics. Reprinted from ref. [32d].

PS templates by sonication led to coalescence. Moreover, the removal of the PS template by calcination was found to damage the TiO₂ shell (Fig. 3A).^[32d] Instead, dissolution of the PS *via* a mild Soxhlet extraction with chloroform kept the shells intact even after a subsequent calcination (Fig. 3B).^[32d] The resulting monodisperse 200 nm anatase TiO₂ NC were used for the photocatalytic degradation of recalcitrant dyes. The photoactivity depended on the shell morphology: broken, holey shells resulting from direct thermal template degradation showed lower activity than intact TiO₂ NC obtained from Soxhlet extraction and subsequent calcination (see Fig. 3D and E). An increase of shell thickness drastically reduced the catalytic activity as well.

Reduction of Ag⁺ by NaBH₄ in presence of the PS template and subsequent TiO₂ coating resulted in monodisperse 200 nm Ag-TiO₂ NC with intact shells before and after calcination.^[32c] Besides their application in photoelectrochemical water splitting,^[38] the Ag-TiO₂ NC were also evaluated for their potential as antimicrobial active material.^[32c] Amorphous as well as anatase TiO₂ NC with 2 or 8 wt% Ag loading were studied for their antimicrobial, immunological and cytotoxic profiles towards murine macrophages and fibroblasts in order to ensure that an eventual antibacterial coating based on these particles would not only be efficient but also biocompatible. Even if the particles are firmly attached to an implant surface, they might detach and thus uptake by immune cells is likely, calling for the investigation of their immunological impact. We found that all studied particles did not induce any immune response and the active uptake of significant amounts, proven by flow cytometry, confocal microscopy, as well as electron microscopy, did not impact the macrophages' viability. A long-term sustained release of Ag⁺ could be measured for all the different Ag-TiO₂ NC, being highest for anatase TiO₂ NC with 8 wt% Ag. In line with that, they provided the highest antimicrobial activity among all samples, being even bactericidal to methicillin resistant *S. aureus*, while showing the same cytocompatibility as the other particles.

Ag@SiO₂ NR Using a Microemulsion Template

Soft templates with flexible shapes are the second possibility for the generation of voids inside nanomaterials. The shells are

deposited *via* sol-gel routes, hydrothermal reactions, emulsion polymerizations or other precipitation reactions at the interface of the soft templates. Emulsions can be especially useful for the synthesis of rattle-type hollow nanostructures as the core particles can be encapsulated or synthesized inside the micelle. In the Fromm group, Ag@SiO₂ NR were prepared in a reverse microemulsion system.^[39] The synthesis exploits the different solubility of two silica precursors to induce a structural difference within the silica shell, which allows for a selective etching of the inner part. The micelle restricts the overall particle size. In the first step, hydrazine reduces Ag⁺ in the aqueous phase to form the Ag cores (Fig. 4A). The shells are then hydrolyzed from tetraethyl orthosilicate (TEOS) and (3-aminopropyl)trimethoxysilane (APTMS). TEOS is better soluble in the cyclohexane phase, whereas APTMS condenses in the aqueous phase around the Ag core. The void between the Ag and the silica shell is formed by short washing in warm water which selectively etches the less 'dense' silica network close to the Ag core derived from APTMS. Variation of the Ag⁺ concentration tunes the ratio of filled Ag@SiO₂ NR to empty silica NC and thus the materials' overall Ag loading (*e.g.* 3, 13 and 23 wt%).^[40] In dynamic release experiments, a continuous Ag⁺ release was observed (enhanced with higher Ag loading), leading to better antimicrobial activity of the Ag@SiO₂ NR with higher Ag content. The particles did not induce inflammation, hemolysis nor did they affect the viability or metabolism of dendritic cells at concentrations that had strong antibacterial effects.^[40] The Ag@SiO₂ NR provided as well good catalytic activity that was studied with the reduction of methylene blue (MB) to its colorless leuco form by NaBH₄ as model reaction.^[39]

Self-templating Strategy for Ag@SiO₂ NR

Self-templating strategies are usually less complex than hard or soft templating as they require fewer synthesis steps. They are reported to have an increased potential for reproducibility, successful scale-up and lower production costs.^[9] Moreover, they avoid the calcination step frequently used for template removal, which is not compatible with temperature-sensitive shell materials, altering shell morphology in an undesired way.^[9,42] Surface

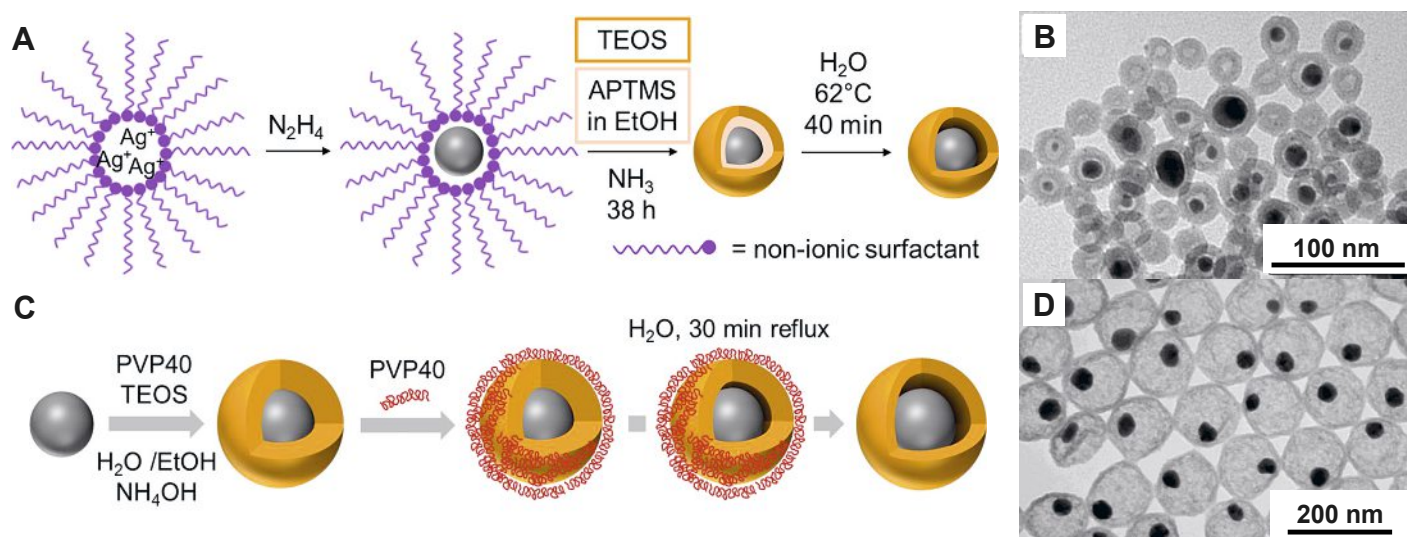


Fig. 4. (A) Microemulsion synthesis of 25 nm Ag@SiO₂ nanorattles and (B) their electron micrograph. Reprinted from ref. [39]. (C) Surface protected etching as self-templating approach to 100 nm Ag@SiO₂ nanorattles and (D) their electron micrograph.[41]

protected etching, as one example of self-templating techniques, enhances the relative stability of a nanoparticle's surface, which enables the selective removal of the interior.^[9,43] In order to realize Ag@SiO₂ NR, solid Ag@SiO₂ core shell (CS) particles were synthesized and coated with a layer of polyvinylpyrrolidone (PVP40) that coordinates strongly with its carbonyl groups to hydroxyl groups on the silica surface *via* hydrogen bonding.^[44] The polymer chains are too large to diffuse inside the shell and thus they protect the particle's outer surface during the etching process. Water is small enough to penetrate the shell and can be used as a mild etching agent on Stöber derived silica, which shows increased water solubility at elevated temperatures due to incomplete polycondensation of the alkoxide precursors.^[45] Our synthesis resulted in very homogeneous Ag@SiO₂ NR with a distinctive void between the Ag core and the thin silica shell, while avoiding empty silica NC.^[41] The protective role of the silica shell was clearly demonstrated again in the catalytic reduction of MB. Contrary to AgNP isolated before the silica coating, Ag@SiO₂ CS and NR led to a fast initial decrease of the MB absorption band and its complete disappearance within 30 min. Furthermore, the Ag@SiO₂ CS and NR showed excellent bactericidal properties, even against methicillin resistant *S. aureus*.^[41] A first promising cytocompatibility evaluation using murine fibroblasts together with the easy way of functionalization using silica precursors with one non-hydrolysable side chain will allow the exploitation of the high antimicrobial efficacy for nanobiomaterials.

Conclusion

Overall, all three synthetic strategies have proven to be useful for the preparation of hollow nanostructured materials, especially for realizing different sizes with hard templates for ≥ 200 nm, self-templating for about 100 nm and the microemulsion for the smallest nanocontainers of 25 nm. While the microemulsion soft template required precise control of the reaction conditions for reproduction and selectivity of the silica nucleation on the Ag cores was hardly achieved, the self-templating approach allowed an easy and fast access to a homogenous encapsulation of the Ag cores. But, in contrast to the hard template approach, both syntheses are rather limited in their yield and batch scale. With regard to the biological evaluation of the Ag-containing particles anatase Ag-TiO₂ NC and the surface protected etching derived Ag@SiO₂ NR were identified as the most promising candidates for introduction into biomaterials, *e.g.* into polymer composites or as coatings on implant surfaces.

Acknowledgements

The authors would like to thank the University of Fribourg and the Swiss National Science Foundation, in particular the National Center of Competence in Research 'Bio-Inspired Materials' for the generous funding of the project.

Received: Nov. 15, 2018

- [1] X. Wang, J. Feng, Y. Bai, Q. Zhang, Y. Yin, *Chem. Rev.* **2016**, *116*, 10983.
- [2] M. Priebe, K. M. Fromm, *Chem. Eur. J.* **2015**, *21*, 3854.
- [3] P. Arosio, R. Ingrassia, P. Cavadini, *Biochim. Biophys. Acta* **2009**, *1790*, 589.
- [4] R. C. Mundargi, M. G. Potroz, S. Park, H. Shirahama, J. H. Lee, J. Seo, N. J. Cho, *Small* **2016**, *12*, 1167.
- [5] T. Trantidou, M. Friddin, Y. Elani, N. J. Brooks, R. V. Law, J. M. Seddon, O. Ces, *ACS Nano* **2017**, *11*, 6549.
- [6] J. Sun, J. Zhang, M. Zhang, M. Antonietti, X. Fu, X. Wang, *Nat. Commun.* **2012**, *3*.
- [7] J. P. Vigneron, M. Rassart, Z. Vértessy, K. Kertész, M. Sarrazin, L. P. Biró, D. Ertz, V. Lousse, *Phys. Rev. E* **2005**, *71*, 011906.
- [8] T. M. Trzeciak, P. Vukusic, *Phys. Rev. E* **2009**, *80*, 061908.
- [9] Q. Zhang, W. Wang, J. Goebel, Y. Yin, *Nano Today* **2009**, *4*, 494.
- [10] J.-H. Lee, *Sens. Actuator B Chem.* **2009**, *140*, 319.
- [11] C. W. Yen, M. A. Mahmoud, M. A. El-Sayed, *J. Phys. Chem. A* **2009**, *113*, 4340.
- [12] J. Liu, S. Z. Qiao, J. S. Chen, X. W. Lou, X. Xing, G. Q. Lu, *Chem. Commun.* **2011**, *47*, 12578.
- [13] a) N. Halas, *MRS Bulletin* **2011**, *30*, 362; b) M. Retsch, M. Schmelzeisen, H.-J. Butt, E. L. Thomas, *Nano Lett.* **2011**, *11*, 1389; c) L. Zhang, H. Wang, *ACS Nano* **2011**, *5*, 3257.
- [14] a) J. Dinesh, M. Uzma, M. Pranab, S. Athinarayanan, E. Muthusamy, *Angew. Chem. Int. Ed.* **2008**, *120*, 7799; b) Y. Wang, Q. Zhu, H. Zhang, *J. Mater. Chem.* **2006**, *16*, 1212.
- [15] C. C. Nguyen, N. N. Vu, T.-O. Do, *J. Mater. Chem., A* **2015**, *3*, 18345.
- [16] a) X. W. Lou, L. A. Archer, Z. Yang, *Adv. Mater.* **2008**, *20*, 3987; b) S. Laifa, Y. Le, Y. Xin-Yao, Z. Xiaogang, L. X. Wen, *Angew. Chem. Int. Ed.* **2015**, *54*, 1868.
- [17] W. Zhiyu, Z. Liang, L. X. Wen, *Adv. Mater.* **2012**, *24*, 1903.
- [18] a) S. S. Kelkar, T. M. Reineke, *Bioconjug. Chem.* **2011**, *22*, 1879; b) J. Xie, S. Lee, X. Chen, *Adv. Drug Delivery Rev.* **2010**, *62*, 1064.
- [19] a) Y. Chen, H. Chen, D. Zeng, Y. Tian, F. Chen, J. Feng, J. Shi, *ACS Nano* **2010**, *4*, 6001; b) W. Huixia, Z. Shengjian, Z. Jiamin, L. Gang, S. Jianlin, Z. Lingxia, C. Xiangzhi, R. Meiling, H. Qianjun, B. Wenbo, *Adv. Funct. Mater.* **2011**, *21*, 1850.
- [20] H. Hu, H. Zhou, J. Du, Z. Wang, L. An, H. Yang, F. Li, H. Wu, S. Yang, *J. Mater. Chem.* **2011**, *21*, 6576.
- [21] X. Yang, S. E. Skrabalak, Z.-Y. Li, Y. Xia, L. V. Wang, *Nano Lett.* **2007**, *7*, 3798.
- [22] a) R. K. Singh, T.-H. Kim, C. Mahapatra, K. D. Patel, H.-W. Kim, *Langmuir* **2015**, *31*, 11344; b) X. Fang, X. Zhao, W. Fang, C. Chen, N. Zheng, *Nanoscale* **2013**, *5*, 2205.
- [23] P. Tanner, P. Baumann, R. Enea, O. Onaca, C. Palivan, W. Meier, *Acc. Chem. Res.* **2011**, *44*, 1039.

- [24] Y. Hu, X. T. Zheng, J. S. Chen, M. Zhou, C. M. Li, X. W. Lou, *J. Mater. Chem.* **2011**, *21*, 8052.
- [25] H. Jiu, W. Jia, L. Zhang, C. Huang, H. Jiao, J. Chang, *J. Porous Mat.* **2015**, *22*, 1511.
- [26] E. Yu, I. Galiana, R. Martínez-Máñez, P. Stroeve, M. D. Marcos, E. Aznar, F. Sancenón, J. R. Murguía, P. Amorós, *Colloids Surf., B* **2015**, *135*, 652.
- [27] a) K. H. Bae, K. Lee, C. Kim, T. G. Park, *Biomaterials* **2011**, *32*, 176; b) M. Xing, Z. Yun, N. K. Woei, Z. Yanli, *Chem. Eur. J.* **2013**, *19*, 15593.
- [28] a) Q. Wang, X. Huang, Y. Long, X. Wang, H. Zhang, R. Zhu, L. Liang, P. Teng, H. Zheng, *Carbon* **2013**, *59*, 192; b) C. Yu, X. Pengfei, W. Meiyang, M. Qingshuo, C. Hangrong, S. Zhu, W. Jin, Z. Lingxia, L. Yaping, S. Jianlin, *Adv. Mater.* **2014**, *26*, 4294.
- [29] K. Patel, S. Angelos, W. R. Dichtel, A. Coskun, Y.-W. Yang, J. I. Zink, J. F. Stoddart, *J. Am. Chem. Soc.* **2008**, *130*, 2382.
- [30] Z. Zhang, F. Wang, Y. S. Sohn, R. Nechushtai, I. Willner, *Adv. Funct. Mater.* **2014**, *24*, 5662.
- [31] F. Caruso, R. A. Caruso, H. Möhwald, *Science* **1998**, *282*, 1111.
- [32] a) J. Gagnon, M. J. D. Clift, D. Vanhecke, I. E. Widner, S. L. Abram, A. Petri-Fink, R. A. Caruso, B. Rothen-Rutishauser, K. M. Fromm, *J. Mater. Chem., B* **2016**, *4*, 1166; b) J. Gagnon, M. J. D. Clift, D. Vanhecke, D. A. Kuhn, P. Weber, A. Petri-Fink, B. Rothen-Rutishauser, K. M. Fromm, *J. Mater. Chem., B* **2015**, *3*, 1760; c) N. Hérault, J. Wagner, S. L. Abram, L. Horváth, J. Widmer, C. Bourquin, K. M. Fromm, *Nanomedicine: NBM* **2019**, submitted; d) N. Hérault, K. M. Fromm, *Helv. Chim. Acta* **2017**, *100*, e1700014.
- [33] L. Zhang, F. Su, X. Kong, F. Lee, K. Day, W. Gao, M. E. Vecera, J. M. Sohr, S. Buizer, Y. Tian, D. R. Meldrum, *RSC Adv.* **2016**, *6*, 46134.
- [34] N. Hérault, J. Gagnon, P. Reñones, F. Harvey, F. Fresno, D. Chen, R. A. Caruso, V. A. de la Peña O'Shea, K. M. Fromm, manuscript under preparation, **2019**.
- [35] D. Chen, X. Zhang, A. F. Lee, *J. Mater. Chem. A* **2015**, *3*, 14487.
- [36] L. Collado, P. Jana, B. Sierra, J. M. Coronado, P. Pizarro, D. P. Serrano, V. A. de la Peña O'Shea, *Chem. Eng. J.* **2013**, *224*, 128.
- [37] Y. Ma, X. Wang, Y. Jia, X. Chen, H. Han, C. Li, *Chem. Rev.* **2014**, *114*, 9987.
- [38] N. Hérault, V. Kaliginedi, P. Moreno, P. Broekmann, K. M. Fromm, manuscript under preparation, **2019**.
- [39] M. Priebe, K. M. Fromm, *Part. Part. Syst. Char.* **2014**, *31*, 645.
- [40] M. Priebe, J. Widmer, N. Suhartha Löwa, S.-L. Abram, I. Mottas, A.-K. Woischnig, P. S. Brunetto, N. Khanna, C. Bourquin, K. M. Fromm, *Nanomedicine: NBM* **2017**, *13*, 11.
- [41] S.-L. Abram, K. M. Fromm, manuscript under preparation, **2019**.
- [42] P. M. Ajayan, O. Stephan, P. Redlich, C. Colliex, *Nature* **1995**, *375*, 564.
- [43] Q. Zhang, T. Zhang, J. Ge, Y. Yin, *Nano Lett.* **2008**, *8*, 2867.
- [44] a) M. A. Cohen Stuart, G. J. Fleer, B. H. Bijsterbosch, *J. Colloid Interface Sci.* **1982**, *90*, 310; b) M. Pattanaik, S. K. Bhaumik, *Mater. Lett.* **2000**, *44*, 352.
- [45] a) L. L. Hench, J. K. West, *Chem. Rev.* **1990**, *90*, 33; b) Y. Hu, Q. Zhang, J. Goebel, T. Zhang, Y. Yin, *Phys. Chem. Chem. Phys.* **2010**, *12*, 11836.